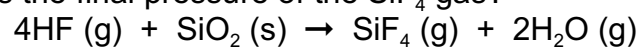


NAME: \_\_\_\_\_

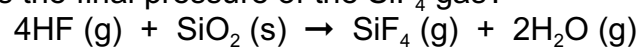
The following reaction was initiated at 122°C at 4.66 atm of HF atm in a constant volume container. The SiO<sub>2</sub> was in excess. At the end of the reaction the temperature was 364°C. What was the final pressure of the SiF<sub>4</sub> gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

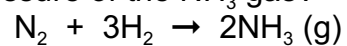
The following reaction was initiated at 170°C at 8.61 atm of HF atm in a constant volume container. The SiO<sub>2</sub> was in excess. At the end of the reaction the temperature was 303°C. What was the final pressure of the SiF<sub>4</sub> gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

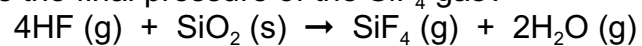
The following reaction was initiated at 160°C at 1.10 atm of H<sub>2</sub> atm in a constant volume container. The N<sub>2</sub> was in excess. At the end of the reaction the temperature was 380°C. What was the final pressure of the NH<sub>3</sub> gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

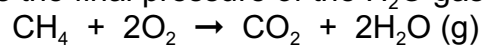
The following reaction was initiated at 284 °C at 7.84 atm of HF atm in a constant volume container. The SiO<sub>2</sub> was in excess. At the end of the reaction the temperature was 472 °C. What was the final pressure of the SiF<sub>4</sub> gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

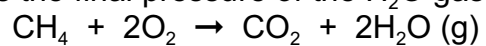
The following reaction was initiated at 254 °C at 7.49 atm of CH<sub>4</sub> atm in a constant volume container. The O<sub>2</sub> was in excess. At the end of the reaction the temperature was 539 °C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

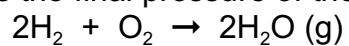
The following reaction was initiated at 218°C at 4.36 atm of CH<sub>4</sub> atm in a constant volume container. The O<sub>2</sub> was in excess. At the end of the reaction the temperature was 514°C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

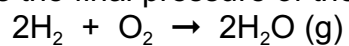
The following reaction was initiated at 199°C at 2.84 atm of O<sub>2</sub> atm in a constant volume container. The H<sub>2</sub> was in excess. At the end of the reaction the temperature was 311°C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

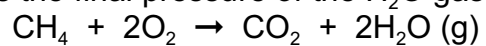
The following reaction was initiated at 100°C at 7.87 atm of O<sub>2</sub> atm in a constant volume container. The H<sub>2</sub> was in excess. At the end of the reaction the temperature was 389°C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

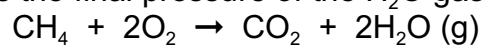
The following reaction was initiated at 271 °C at 3.37 atm of CH<sub>4</sub> atm in a constant volume container. The O<sub>2</sub> was in excess. At the end of the reaction the temperature was 537 °C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

The following reaction was initiated at 252°C at 4.64 atm of CH<sub>4</sub> atm in a constant volume container. The O<sub>2</sub> was in excess. At the end of the reaction the temperature was 520°C. What was the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

## Homework question 16 – Stoichiometry

copy 111 1.88\*

copy 112 2.80\*

copy 113 1.11\*

copy 114 2.62\*

copy 115 23.08\*

copy 116 13.98\*

copy 117 7.03\*

copy 118 27.93\*

copy 119 10.03\*

copy 120 14.02\*

\*Note the sig figs are not necessarily correct. The program cannot figure them.