Homework question 16 – Stoichiometry copy 111

NAME:_____

The following reaction was initiated at 122°C at 4.66 atm of HF atm in a constant volume container. The SiO_2 was in excess. At the end of the reaction the temperature was 364°C. What was the final pressure of the SiF_4 gas?

$4HF\ (g) + SiO_2\ (s) \rightarrow SiF_4\ (g) + 2H_2O\ (g)$

ANS:_____

Homework question 16 – Stoichiometry copy 112

NAME:_____

The following reaction was initiated at 170°C at 8.61 atm of HF atm in a constant volume container. The SiO_2 was in excess. At the end of the reaction the temperature was 303°C. What was the final pressure of the SiF_4 gas?

$4HF\ (g) + SiO_2\ (s) \rightarrow SiF_4\ (g) + 2H_2O\ (g)$

ANS:_____

Homework question 16 – Stoichiometry copy 113

NAME:_____

The following reaction was initiated at 160°C at 1.10 atm of H_2 atm in a constant volume container. The N_2 was in excess. At the end of the reaction the temperature was 380°C. What was the final pressure of the NH_3 gas?

$N_2 + 3H_2 \rightarrow 2NH_3$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 114

NAME:_____

The following reaction was initiated at 284°C at 7.84 atm of HF atm in a constant volume container. The SiO_2 was in excess. At the end of the reaction the temperature was 472°C. What was the final pressure of the SiF_4 gas?

$4HF\ (g) + SiO_2\ (s) \rightarrow SiF_4\ (g) + 2H_2O\ (g)$

ANS:_____

Homework question 16 – Stoichiometry copy 115

NAME:_____

The following reaction was initiated at 254°C at 7.49 atm of CH_4 atm in a constant volume container. The O_2 was in excess. At the end of the reaction the temperature was 539°C. What was the final pressure of the H_2O gas?

$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 116

NAME:_____

The following reaction was initiated at 218°C at 4.36 atm of CH_4 atm in a constant volume container. The O_2 was in excess. At the end of the reaction the temperature was 514°C. What was the final pressure of the H_2O gas?

$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 117

NAME:_____

The following reaction was initiated at 199°C at 2.84 atm of O_2 atm in a constant volume container. The H_2 was in excess. At the end of the reaction the temperature was 311°C. What was the final pressure of the H_2O gas?

$2H_2 + O_2 \rightarrow 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 118

NAME:_____

The following reaction was initiated at 100°C at 7.87 atm of O_2 atm in a constant volume container. The H_2 was in excess. At the end of the reaction the temperature was 389°C. What was the final pressure of the H_2O gas?

$2H_2 + O_2 \rightarrow 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 119

NAME:_____

The following reaction was initiated at 271°C at 3.37 atm of CH_4 atm in a constant volume container. The O_2 was in excess. At the end of the reaction the temperature was 537°C. What was the final pressure of the H_2O gas?

$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry copy 120

NAME:_____

The following reaction was initiated at 252°C at 4.64 atm of CH_4 atm in a constant volume container. The O_2 was in excess. At the end of the reaction the temperature was 520°C. What was the final pressure of the H_2O gas?

$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ (g)

ANS:_____

Homework question 16 – Stoichiometry

copy 111 1.88*

copy 112 2.80*

copy 113 1.11*

copy 114 2.62*

copy 115 23.08*

copy 116 13.98*

copy 117 7.03*

copy 118 27.93*

copy 119 10.03*

copy 120 14.02*

*Note the sig figs are not necessarily correct. The program cannot figure them.