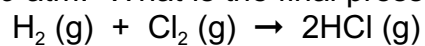


NAME: \_\_\_\_\_

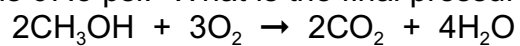
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $\text{Cl}_2$  gas is reacted with an excess of  $\text{H}_2$ . The starting pressure for the  $\text{Cl}_2$  is 77.9 atm. What is the final pressure of the  $\text{HCl}$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

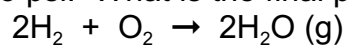
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $O_2$  gas is reacted with an excess of  $CH_3OH$ . The starting pressure for the  $O_2$  is 67.5 psi. What is the final pressure of the  $H_2O$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

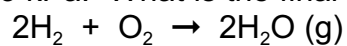
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $\text{O}_2$  gas is reacted with an excess of  $\text{H}_2$ . The starting pressure for the  $\text{O}_2$  is 82.5 psi. What is the final pressure of the  $\text{H}_2\text{O}$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

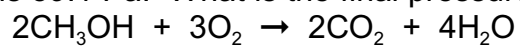
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $\text{O}_2$  gas is reacted with an excess of  $\text{H}_2$ . The starting pressure for the  $\text{O}_2$  is 15.3 kPa. What is the final pressure of the  $\text{H}_2\text{O}$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

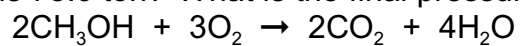
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature. O<sub>2</sub> gas is reacted with an excess of CH<sub>3</sub>OH . The starting pressure for the O<sub>2</sub> is 59.1 Pa. What is the final pressure of the H<sub>2</sub>O gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

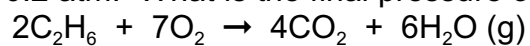
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $O_2$  gas is reacted with an excess of  $CH_3OH$ . The starting pressure for the  $O_2$  is 78.6 torr. What is the final pressure of the  $CO_2$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

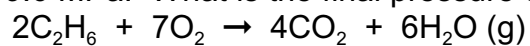
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $O_2$  gas is reacted with an excess of  $C_2H_6$ . The starting pressure for the  $O_2$  is 70.2 atm. What is the final pressure of the  $H_2O$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

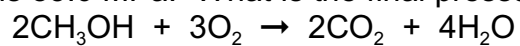
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $\text{O}_2$  gas is reacted with an excess of  $\text{C}_2\text{H}_6$ . The starting pressure for the  $\text{O}_2$  is 39.6 MPa. What is the final pressure of the  $\text{H}_2\text{O}$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

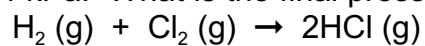
The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $O_2$  gas is reacted with an excess of  $CH_3OH$ . The starting pressure for the  $O_2$  is 83.6 MPa. What is the final pressure of the  $H_2O$  gas?



ANS: \_\_\_\_\_

NAME: \_\_\_\_\_

The following reaction was performed in a rigid volume and the temperature was returned to the starting temperature.  $\text{Cl}_2$  gas is reacted with an excess of  $\text{H}_2$ . The starting pressure for the  $\text{Cl}_2$  is 15.4 kPa. What is the final pressure of the HCl gas?



ANS: \_\_\_\_\_

## Homework question 15 – Stoichiometry

copy 111 155.80 atm.\*

copy 112 90.00 psi.\*

copy 113 165.00 psi.\*

copy 114 30.60 kPa.\*

copy 115 78.80 Pa.\*

copy 116 52.40 torr.\*

copy 117 60.17 atm.\*

copy 118 33.94 MPa.\*

copy 119 111.47 MPa.\*

copy 120 30.80 kPa.\*

\*Note the sig figs are not necessarily correct. The program cannot figure them.